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each isotope as in Example 10.2 and
Answer 10.8. ${}_{9}^{20}\text{F} + {}_{-1}^{0}\text{e} \rightarrow {}_{10}^{20}\text{Ne}$ a. ${}_{38}^{92}\text{Sr} + {}_{-1}^{0}\text{e} \rightarrow {}_{39}^{92}\text{Y}$ b. ${}_{24}^{55}\text{Cr} + {}_{-1}^{0}\text{e} \rightarrow {}_{25}^{55}\text{Mn}$ 10.10 Write a
balanced nuclear equation for positron
emission as in Example 10.3. a. [1] Write
an incomplete equation with the original
nucleus on the left and the particle

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Assessment page 372 5. Compare the mass of the reactants and the mass of the products in a chemical reaction, and explain how these masses are related. The coefficients in the balanced equation indicate the molar relationship

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continued Write the formula for each of the following acids in the space provided. 22. hydrosulfuric acid 23. nitric acid 24. phosphorous acid 25. perchloric acid Refer to the equation below to answer questions 26 and 27.

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nonpolar covalent bond. When the atoms in a bond pull equally and the bonding electrons are shared equally.
polar covalent bond.

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nuclear chemistry from dis- ... Section 25.1 Nuclear Radiation ... iodine-131 into xenon-131 by beta-particle emission, as shown in Figure 25-5. Note that the mass number of the product nucleus is the same

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CHAPTER 4 SOLUTIONS MANUAL Section 4.1 Early Ideas About Matter pages 102–105 Section 4.1 Assessment page 105 1. Contrast the methods used by the Greek philosophers and Dalton to study the atom. Greek philosophers could not

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