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water,  $K_w$  is  $(1.0 \times 10^{-14})$  (at  $25^\circ \text{C}$ ), the  $(pK_w)$  is 14, the constant of water determines the range of the pH scale. To understand what the  $pK_w$  is, it is important to understand first what the "p" means in  $pOH$  and  $pH$ . The addition of the "p" reflects the negative of the logarithm,  $(-\log)$ . Therefore, the  $pH$  is the negative logarithm of the ...

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